THE OXIDATION BEHAVIOR OF YTTERBIUM SCANDIUM ZIRCONIUM NITRIDES UNDER NONISOTHERMAL CONDITIONS

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ABSTRACT

Ytterbium codoped scandium stabilized zirconia, $\text{Zr}_{0.60}\text{Sc}_{0.175}\text{Yb}_{0.225}\text{O}_{1.8}$ was synthesized and nitried for 48 h under an ammonia atmosphere. The compound obtained was nominally $\text{Zr}_{0.60}\text{Sc}_{0.175}\text{Yb}_{0.225}\text{O}_{0.0627}\text{N}_{1.157}$ with a red color and a cation deficient rocksalt structure. Its stability and reoxidation behavior was studied by differential thermal analysis (DTA) and thermogravimetric analysis (TG) under non-isothermal conditions. For evaluation of the experimental data, the Coats Redfern method [1] was used and revealed that the decomposition of the nitride to form ytterbium scandium zirconium oxide is diffusion controlled. The activation energy for the decomposition process was found to be 506 KJ/mol compared to 103 KJ/mol for $\text{ZrN}$ [2]. The results indicate that scandium-zirconium-nitrides co-doped with lanthanide metals undergo a reduction in $\text{O}^2-$ and $\text{N}^3-$ mobility in the crystal and consequently a delay in the reoxidation process.

Keywords: Thermal stability of nitrides, non-isothermal TG/DTA/DDTA, ytterbium, scandium zirconium.
1. INTRODUCTION

Zirconium nitrides have lately gained considerable interest because of their properties such as super hardness, electronic conductivity and improved thermal and mechanical behavior. These nitrides are widely used as wear-resisting coatings, structural materials and diffusion barriers in advanced device semiconductor technologies [3].

Among these metal nitrides, ZrN and scandium stabilized zirconium nitride (ScSZN), are potential materials for future applications in microelectronics such as thin film resistors and barrier ceramics. These nitride materials are best prepared by a nitridation using ammonia at around 1473 K. During this process the corresponding oxynitrides and finally nitrides are formed.

The oxynitrides were found to crystallize sometimes with a fluorite related superstructure but the nitrides crystallize with a cation deficient rock salt structure. These latter materials then have quite reasonable electronic and ionic conductivities [4,5,6]. An understanding of the reoxidation behavior on air as well as the thermal stability of a nitride like YbScSZN is important, if such materials are used for coatings.

From the literature, only a few models for the reaction path of such reoxidation reactions are known. In the case of the reoxidation of Si3N4 several processes are said to be involved and it is assumed that the rate constant might be limited by oxygen diffusion, nitrogen diffusion by both or additionally by a reaction including an interface oxygen diffusion, nitrogen diffusion by both or finally the desorption of N2.

The activation energy for nitrogen diffusion in Sc and C dopant ScSZN stabilized zirconia was found to be 1.8 eV (174 kJ/mol) while for oxygen was 1eV (96 kJ/mol) [9]. This means that oxygen diffusion starts first. The presence of an additional lanthanide ion in scandia stabilized zirconia will reduce the ionic and electronic conductivity in such compounds.

2. EXPERIMENTAL PART

A microcrystalline powder of Zr0.60Sc0.175Yb0.225O1.8 (YbScSZ) with a molar ratio Zr/Sc 48/14 containing 3 mol% of Yb2O3 was synthesized by the polymer complex method [10]. This compound was characterized using X-ray diffraction (XRD), scanning electron microscopy (SEM) and energy dispersive X-ray analysis (EDX), and then reacted for 2d at 1473 K with ammonia to generate Zr0.60Sc0.175Yb0.225O0.0627N1.157 with a red color and a cation deficient sodium chloride structure.

The different phases were identified by XRD within a scan range between 10° - 110° (XRD; Siemens D5000 0/0, CuKλ (λ = 1.5418 Å)).

Thermogravimetric and differential thermal analysis (TG/DTA) experiments were carried out using a Netzsch STA 449F3 Thermoanalyser. Scans were measured in an atmosphere of air at different heating rates (2, 5 and 10 K/min).

3. RESULTS AND DISCUSSION

Rietveld refinement of the X-ray diffractogram of Zr0.60Sc0.175Yb0.225O1.8 established R3 as space group for the fluorite related supercell with lattice constants a = 9.5399(7) Å and c = 17.4770(8) Å, as well as a random distribution of the cations Zr/Sc/Yb. After nitridation, the diffractogram of the corresponding nitride, Zr0.60Sc0.175Yb0.225O0.0627N1.157 exhibited a cation deficient rock salt structure Fm 3m with a cell parameter a = 4.5431(1) Å. These cells were larger compared to Zr0.60Sc0.26O2.63 [5] (R3 : a = 9.4941(1) and c = 17.4252(1) Å) due to the difference of ionic radius between Sc3+ (87 pm) and Yb3+ (98 pm) [11]. The results of SEM/EDX analysis for Zr0.60Sc0.175Yb0.225O1.8 proved Yb and Sc as dopant cations (YbM5 = 1.526 keV and ScK = 4.493 keV).
3.1 TG and DTA results

The sample weight increased continuously up to 1100 K during the reoxidation reaction. The mayor weight increase (90%) occurred between 850 and 1010 K. Analyses of the DTA curves showed that the increase took place in two exothermic steps for YbScZN while only one exothermic step was found for ScZN.

The reduction process of oxygen:

\[ \text{O}_2 + 4 \text{e}^- \rightarrow 2 \text{O}^{2-} \]

is endothermic and requires the dissociation energy of \( \text{O}_2 \) (498 kJ/mol) as well as the electron affinity (704 kJ/mol) to sum 1202 kJ/mol.

The oxidation process of \( \text{N}^{3-} \):

\[ 2\text{N}^{3-} \rightarrow \text{N}_2 + 6 \text{e}^- \]

is exothermic and calculated from the association energy (-945 kJ/mol) and the negative electron affinity (N\(^{3-} \rightarrow \text{N} \) -2300 kJ/mol) to be -3245 kJ/mol. So this redox reaction:

\[ 4\text{N}^{3-} + 3\text{O}_2 \rightarrow 6\text{O}^{2-} + 2\text{N}_2 \]

will produce at the end one strong exothermic signal in the DTA measurements.

However, if \( \text{N}_2 \) bubbles are formed at an interface between the nitride and an \( \text{O}_2 \) saturated reoxidized surface layer, the release of dinitrogen might be delayed. The dissolution of \( \text{N}_2 \) gas bubbles then will release the surface energies of the bubbles resulting in a second exothermic peak in the DTA measurements.

Fig. 1 shows the TG/DTA data (heating rate 2K/min) for the reoxidation of ScSZN and it can be observed a strong exothermic signal at 838 K. In this case, both the reoxidation and the \( \text{N}_2 \) release occurs simultaneously because nitrogen diffusion is not hindered and an interface not formed. This is supposed to be different in YbScSZN. While \( \text{Zr}^{4+} \) and \( \text{Sc}^{3+} \) (84 pm, 87 pm) have similar ionic radio. The replacement of \( \text{Zr}^{4+} \) or \( \text{Sc}^{3+} \) by \( \text{Yb}^{3+} \) (98 pm) [11] will lower the ionic mobility of \( \text{N}^{3-} \) and \( \text{O}^{2-} \) as well as the diffusion of \( \text{N}_2 \) and \( \text{O}_2 \).

Figs. 2-4 show the TG/DTA/DDTA curves for the reoxidation of YbScSZN between 320 – 1200 K at different heating rates. In order to compare the results in each run, a sample of 65 mg of YbScSZN was used. At every heating rate the reoxidation reaction starts between 800 and 850 K. Contrary to the results for ScSZN (fig. 1), two different exothermic processes are clearly observed. This might be due to remarkable slower dinitrogen diffusion as a result of the formation of \( \text{N}_2 \) bubbles at the interface.

For the heating rate of 2 K/min (fig. 2) the complete oxidation lasts 45 min and occurs between 835 - 920 K. For 5 K/min the time is 23 min and arises between 870 - 930 K while for a heating rate of 10 K/min the overall reaction time is 20 min and it takes place between 851 - 1057 K. This means the higher the heating rate the more separated are the two exothermic peaks. In addition, while more \( \text{N}_2 \) bubbles are evolved, the interface become thicker, consequently the second exothermic peak appears.
stronger. The complete O\textsubscript{2} uptake and the redox reaction occur at higher heating rates in a shorter time and a longer temperature interval (851-1057 K) and the N\textsubscript{2} release occurs at higher temperature with higher intensity. A deconvolution of the three different DTA-curves gives the results showed in table 1.

Table 1. Data of the two exothermic peaks after deconvolution for the DTA curve

<table>
<thead>
<tr>
<th>Heating rate K/min</th>
<th>Redox reaction Curve maximum at Temperature K</th>
<th>Curve Area</th>
<th>N\textsubscript{2}-release Curve maximum at Temperature K</th>
<th>Curve Area</th>
</tr>
</thead>
<tbody>
<tr>
<td>2</td>
<td>580.8</td>
<td>116.5</td>
<td>629.9</td>
<td>4.8</td>
</tr>
<tr>
<td>5</td>
<td>589.8</td>
<td>193.8</td>
<td>647.3</td>
<td>92.3</td>
</tr>
<tr>
<td>10</td>
<td>598.1</td>
<td>510.4</td>
<td>730.5</td>
<td>349.8</td>
</tr>
</tbody>
</table>

3.2 Crystallization, mechanism and activation energy

The original mass loss versus temperature (TG) curves obtained at constant heating rates were transformed into the degree of conversion (α) versus temperature by 

\[
\alpha = \frac{m_t - m_f}{m_o - m_f},
\]

where \(m_t\) represents the mass of the sample at arbitrary time \(t\) (or temperature \(T\)), whereas \(m_o\) and \(m_f\) are the mass of the sample at the beginning and at the end of the process, respectively (fig. 5).

According to Gualtieri et al. [12] the rate of a kinetically controlled process can be expressed as:

\[
\frac{d\alpha}{dt} = k f(\alpha) \quad \text{or in the integral form } g(\alpha) = kt
\]

(1)

Where \(\alpha\) is the conversion factor, \(k\) the rate constant with respect to time \(t\), \(f(\alpha)\) is the kinetic model function which depends on the reactions mechanism and \(g(\alpha)\) is the integral of \(1/f(\alpha)\).

A non isothermal experiment takes place under a constant temperature increase. \(T_0\) is the temperature at the beginning of the reaction and \(b\) the rate of temperature increase. The temperature at a time \(t\) is
then given by:

\[ T = bt + T_0 \]  

(2)

In order to extract kinetic data from these non-isothermal measurements, it was employed the method developed by Coats and Redfern [1]. It is based on the equation

\[ \ln \frac{g(\alpha)}{A} = \alpha \ln T - \ln \left( \frac{EA}{Eb} \right) - \frac{E}{RT} \]  

(3)

Parameters A and E are determined using several expressions reported in table 2.

On the other hand, the method proposed by Kennedy and Clark [13] is given by:

\[ b^*g(\alpha) = AE^{-\frac{E}{RT}} \]

Taking the logarithm of both sides of this equation gives:

\[ \ln \left( \frac{b^*g(\alpha)}{T-T_0} \right) = \ln(A) - \frac{E}{RT} \]  

(5)

Plotting the left hand side of this equation against 1/T should give a straight line with gradient –E/RT and an intercept of ln (A). Expressions for g(\alpha) are in the Table 2.

**Table 2.** Different expressions for g(\alpha) according to different kinetic models.

<table>
<thead>
<tr>
<th>Kinetic model</th>
<th>g(\alpha)</th>
</tr>
</thead>
<tbody>
<tr>
<td>John Mehl Avrami</td>
<td>([-\ln(1-\alpha)^{1/n}])</td>
</tr>
<tr>
<td>Phase boundary controlled reaction (contracting area)</td>
<td>([1 - (1-\alpha)^{1/2}])</td>
</tr>
<tr>
<td>Phase boundary controlled reaction (contracting volume)</td>
<td>([1 - (1-\alpha)^{1/3}])</td>
</tr>
<tr>
<td>One dimensional diffusion</td>
<td>O(\alpha^2)</td>
</tr>
<tr>
<td>Two dimensional diffusion (Bidimensional particles shape)</td>
<td>((1-\alpha)\ln(1-\alpha) + \alpha)</td>
</tr>
<tr>
<td>Three dimensional diffusion (Tridimensional particles shape)</td>
<td>([1 - (1-\alpha)^{1/3}]^2)</td>
</tr>
<tr>
<td>Jander Equation</td>
<td>([1 - (2\alpha/3) - (1-\alpha)^{2/3}])</td>
</tr>
<tr>
<td>Ginstling-Brounshtein</td>
<td>([1 - (2/3)*(1-\alpha)] - [(1-\alpha)^{2/3}])</td>
</tr>
</tbody>
</table>

In this context, figure 6 shows the reoxidation behavior of Zr_{0.60}Sc_{0.175}Yb_{0.225}O_{0.0627}N_{1.157} as a plot of g(\alpha) versus 1/T for three different heating rates namely 2, 5 and 10 K/min, under dry air and normal atmospheric pressure. The curve is interpreted in four steps.

In the initial and final stage, the oxidation proceeds approximately linear with time, in a temperature range 300 to 500 K and above 1200 K. The linear behavior below 500 K might be expressed by

\[ g(\alpha) = K_1t \]  

(4)

where \(K_1\) is the linear rate constant and t the reaction time. In the temperature range 300 - 500 K the cation deficient nitride with NaCl structure is still present. In step II (temperature range 600 - 700 K) the oxidation process will slow down whereas in step III (above 700 K) oxidation will proceed faster. These steps can be described using the Ginstling-Brounstein equation [14]

\[ g(\alpha) = [1 - (2/3)*(1-\alpha)] - [(1-\alpha)^{2/3}] = K_{j1} * t \]  

(5)

where \(K_{j1}\) is the diffusion-controlled rate constant. In the step II, the degree of oxidation is up to 90% and is characteristic of an autocatalytic reaction behavior. Step III can be related to a diffusion controlled mechanism and it is attributed to the crystallization of the doped zirconium oxide.

\[ g(\alpha) = \alpha^2 = K_{j2} * t \]  

(6)

where \(K_{j2}\) is the diffusion-controlled rate constant in one dimension.

Fig. 6 shows the curves fitting using the equations of Coats and Redfern [1], Kennedy and Clark [3] and equations for the three dimensional diffusion mechanism from Ginstling-Brounstein [14] between 700 and 1000 K. For the three heating rates (2, 5 and 10 K/min), it was observed an excellent linear correlation.

The autocatalytic reaction proceeds up to 980 K with a 90% weight increase. Throughout this process, the cell volume decreases, molecular nitrogen is lost and the oxynitride is formed. Simultaneously, it is produced a diffusion barrier for dinitrogen, consisting of oxide and oxynitride.
The autocatalytic process may originate from interfacial differences between the educt and the product phase. Nevertheless, the N₂ gas developed during reaction and the formation of gas bubbles play a major role in the reactions mechanism. Fig. 7 shows g(α) versus 1000/T /K for the completion of the oxidation reaction with the formation of the fluorite structure using the g(α) = (α)² expression for an one dimensional diffusion-controlled mechanism. From the slope of these lines, the activation energies were calculated.

### 4. CONCLUSIONS

The transformation of YbScSZN to a crystalline oxide phase was studied using TG/DTA experiments, under non-isothermal conditions in order to extract kinetic data. The activation energy \( E_a = 506 \text{ KJ/mol} \) was determined using the Coats Redfern equation and different reaction mechanisms. The best fitting for this oxidation process was achieved using the expressions g(α) corresponding to the one dimensional and three dimensional diffusion controlled mechanisms.

\[
g (\alpha) = [\alpha]^2
\]

\[
g (\alpha) = [1- [(2/3)*(1-\alpha)] - [(1-\alpha)^{(2/3)}]] = K_j \cdot t
\]  

This study suggests that the oxidation kinetics follow two different paths: oxygen diffusion and reaction with nitrogen followed by a slow evolution of molecular nitrogen. The present work shows that the oxidation of the nitride is a very complex diffusion controlled process.

### 5. ACKNOWLEDGEMENTS

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### 6. REFERENCES

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